Lab notes:

10-1-2020

Question: How can we use the law of defining proportions to

Hypostasis: we can decompose water into hydrogen and oxygen, should get 2 to 1 ratio

Equation: H2O - H2 +O2

2H2O – 2H2 +O2

* Current through water
* Fill complete with water
* Gas bubble
  + Different size of gas bubble
* Hydrogen produces gas faster while oxygen is much slower
* REMEMBER OXYGEN HAD HEADSTART BIG BUBBLE IN BERFORE TEST
* Going to burn hydrogen with oxygen which will be
  + 2H2 + 02 – 2H2O

2Hcl +22n – 22ncl + H2

6:0 Hydrogen: oxygen == 0 Stars (nothing happened, no oxygen to make a explosion with fire)

5:1 Hydrogen: oxygen == 2 stars (it went of the lighter because there was oxygen and good fuel)

4:2 Hydrogen: oxygen == 3 Stars (was decent, but Mr. Hansen said 3 stars)

3:3 Hydrogen: oxygen == 3 stars (“Shouldn’t be good but to be fair to science”)

2:4 Hydrogen: oxygen == 2 Stars

1:5 Hydrogen: oxygen == 0 stars

0:6 Hydrogen: oxygen == 0 Stars

Tubes at end of class is not exactly 2:1 more like 4:1

Maybe iron Hydroxide sits at the bottom of test tube that is holding the two tubes and water

In conclusion it did not prove that it is 2:1 but it did prove that (something I did not hear)

Star rating conclusion did work and had good results